

Inorganic Chemistry 1
 CHEMICAL BONDING
 LEWIS STRUCTURES
 Review Questions

Q₁: Why is $\text{:}\ddot{\text{N}}-\ddot{\text{N}}\text{:}$ an incorrect structure for N₂?

Q₂: Write Lewis structures for the following:

- NO⁺ (nitrosyl ion)
- COCl₂ (carbonyl chloride)

Q₃: Write a Lewis structure for each of the following: [One central Atom]

- SOCl₂
- OF₂
- H₂S

Q₄: Write a Lewis structure for each of the following [More than one central atom]

- C₂H₆O; no O-H bonds - dimethyl ether
- NH₃O - hydroxylamine.

Q₅: Write a Lewis structure for each of the following: [Multiple bonds - double/triple bonds]

- CO₂
- CO - only compound where C has triple bond.

Q₆: Draw a Lewis structure for:

- PH₄⁺
- C₂F₄
- SbH₃
- COF₂ (C is central)
- CH₄S
- NO₂⁺
- NO₂F (N is central)
- HNO₃ (HONO₂)
- S₂Cl₂

Q₇: Which are following are correct Lewis Structures and which are not? Explain what is wrong with the incorrect structure.

