

Write the ground-state electronic configuration for

a) Phosphorous, P atom ($Z = 15$) and

b) Ferum, Fe atom ($Z = 26$)

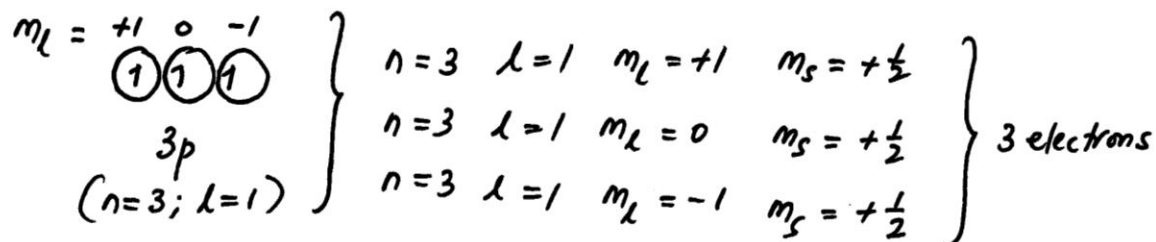
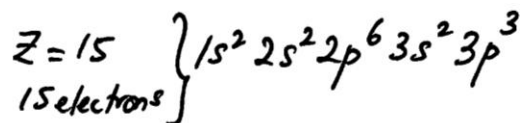
How many unpaired electron are there in these atoms .
Specify all the quantum number (n, l, m_l and m_s) for the unpaired electrons .

Solution

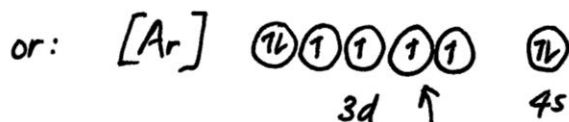
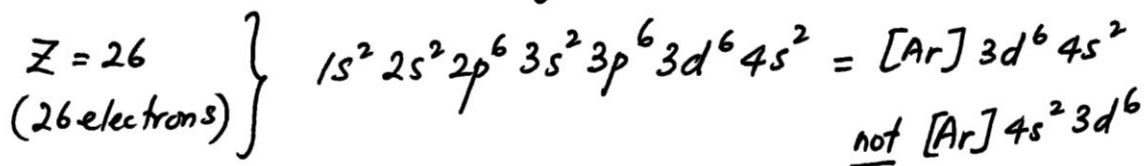


Solution :

a) Ground-state electronic configuration of P atom is :



b) Ground-state electronic configuration of Fe atom is :



4 unpaired electrons

