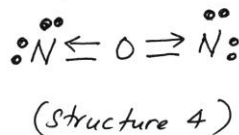
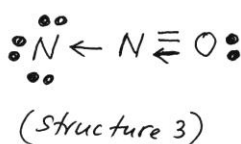
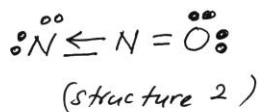
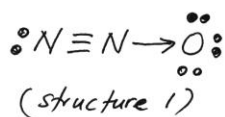


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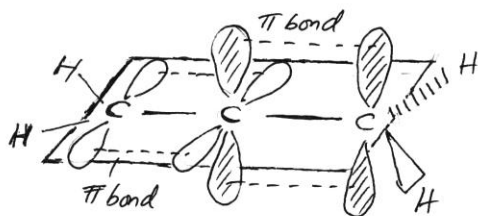
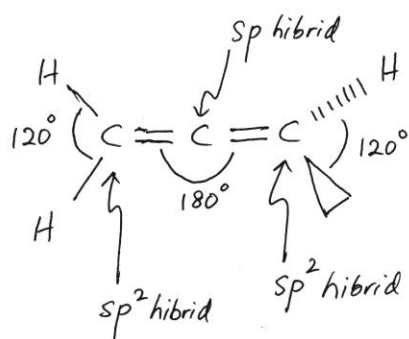
Q5 (a) Lewis structure of N_2O



Given: $N-N$ (bond length = 1.13 \AA) } is for the triple
 $N-O$ (bond length = 1.19 \AA) } bond of $N\equiv N$
 and $N\equiv O$

Therefore structure (1) and structure (3) (both have triple bond) contribute towards the resonance hybrid.

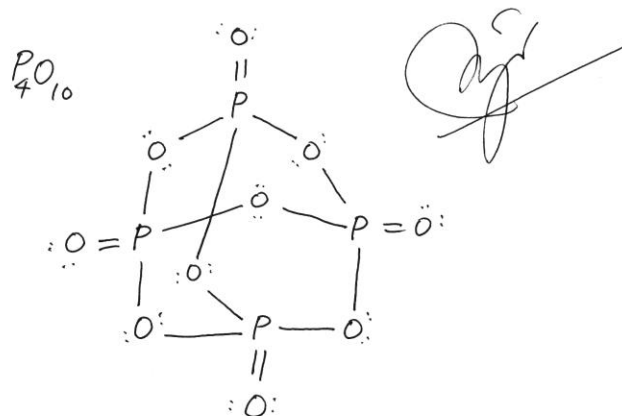
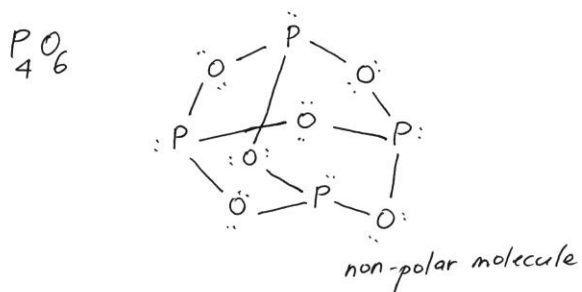
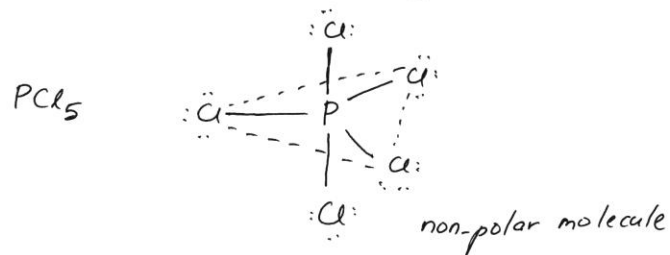
Q5 (b)



$\frac{1}{2}$

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Q5 (c)



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