Inorganic Chemistry 1 QUANTUM THEORY WAVELENGTH : FREQUENCY : WAVE-NUMBER

 $c = \lambda v$

 $C = \text{speed of light} \left(2.998 \times 10^{\text{ms}^{-1}}\right)$ $\lambda = \text{wavelength} \left(\text{meter,m}\right) - \text{lamda}$ $\mathcal{V} = \text{niu} \left(\text{frequency}\right) : \text{Hz} @ \text{s}^{-1}$

l picometer, pm=10¹²m l nanometer, nm=10⁹m | micrometer, $\mu m = 10^6 m$ | angstroms, $A = 10^{-10} m$ 1 Hz = 15"

Converting Frequency, Wavelength and Wave-number **Problem-solving Example 1**

A new laser light has a wavelength of 485 Angstroms. What is the frequency of this light?

$$\lambda = (485 \times 10^{-10})m$$

$$C = 2.998 \times 10^{8} ms^{-1}$$

$$\mathcal{V} = ? Hz$$

$$C = \lambda \mathcal{V} \text{ or } \mathcal{V} = \frac{C}{\lambda} = \frac{2.998 \times 10^{8} ms^{-1}}{485 \times 10^{-10} m}$$

$$= 6.18 \times 10^{+15} Hz$$

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